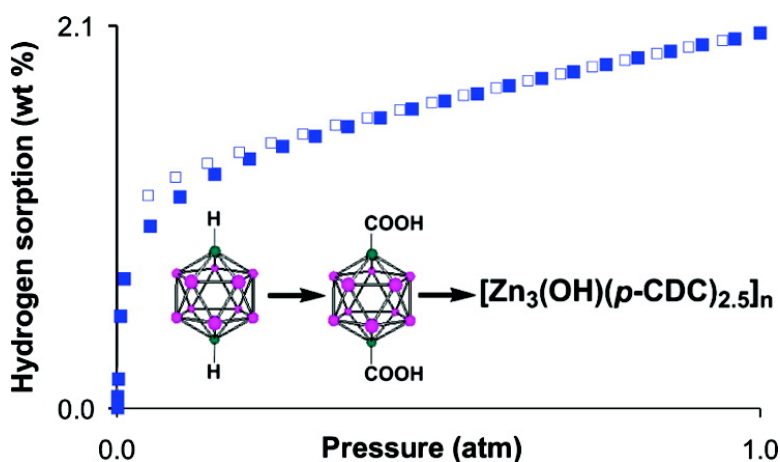


Synthesis and Hydrogen Sorption Properties of Carborane Based Metal–Organic Framework Materials

Omar K. Farha, Alexander M. Spokoyny, Karen L. Mulfort, M. Frederick Hawthorne, Chad A. Mirkin, and Joseph T. Hupp

J. Am. Chem. Soc., **2007**, 129 (42), 12680-12681 • DOI: 10.1021/ja076167a • Publication Date (Web): 27 September 2007

Downloaded from <http://pubs.acs.org> on February 14, 2009



More About This Article

Additional resources and features associated with this article are available within the HTML version:

- Supporting Information
- Links to the 18 articles that cite this article, as of the time of this article download
- Access to high resolution figures
- Links to articles and content related to this article
- Copyright permission to reproduce figures and/or text from this article

[View the Full Text HTML](#)

Synthesis and Hydrogen Sorption Properties of Carborane Based Metal–Organic Framework Materials

Omar K. Farha,[†] Alexander M. Spokoyny,[†] Karen L. Mulfort,^{†,‡} M. Frederick Hawthorne,[§] Chad A. Mirkin,^{*,†} and Joseph T. Hupp^{*,†}

Department of Chemistry and International Institute for Nanotechnology, Northwestern University, 2145 Sheridan Road, Evanston, Illinois 60208, Chemistry Division, Argonne National Laboratory, Argonne, Illinois 60439, and International Institute of Nano and Molecular Medicine, University of Missouri-Columbia, Missouri 65212

Received August 15, 2007; E-mail: j-hupp@northwestern.edu; chadnano@northwestern.edu

Tailorable inorganic coordination polymers,^{1–7} in particular, metal–organic frameworks (MOFs)^{2–7} comprise an important emerging class of materials. They are noteworthy for their structural and chemical diversity, high internal surface areas, and (often) permanent microporosity. As such, they have attracted great interest for numerous applications including ion exchange,² heterogeneous catalysis,³ optoelectronics,⁴ gas separation,⁵ gas sensing,⁶ and gas storage, in particular H₂ storage.⁷ Among the factors useful for attaining high gravimetric uptake of H₂ are small pores, open metal coordination sites, and low framework mass. Here we report on the synthesis and hydrogen uptake properties of carborane-based MOFs that exploit some or all of these characteristics to achieve surprisingly high H₂ uptake. To our knowledge, these are the first metal–organic frameworks to make use of boron-rich components as building blocks.⁸ Carboranes are icosahedral carbon-containing boron clusters possessing several material-favorable properties including rigidity, thermal stability, and chemical stability. Dicarboranes of the form C₂B_{*n*}–2H_{*n*} (6 ≤ *n* ≤ 12) may be regarded as three-dimensional delocalized aromatic systems in which surface bonding and core bonding correspond to σ-bonding and π-bonding, respectively.⁹ These compounds can be prepared on the kilogram scale and have been used for a variety of applications, including boron neutron capture therapy, molecular delivery devices in biomedicine, and molecular motors.¹⁰

As a starting point for carborane-based MOFs, we have focused on Zn(II) coordination of the deprotonated form of 1,12-dihydroxy-carbonyl-1,12-dicarba-*closo*-dodecaborane (*p*-CDCH₂ (**1**); Figure 1A). *p*-CDC^{2–} has approximately the same 2-D footprint as benzene-1,4-dicarboxylate (bdc), the strut defining the archetypal cubic framework compound, MOF-5 [Zn₄O(bdc)₃]_{*n*}. In comparison with other frameworks, MOF-5 is a good, albeit not spectacular, hydrogen storage material, at least at cryogenic temperatures (1.25 wt % at 77 K and 1 atm).¹¹ We reasoned that by replacing benzene with the three-dimensional carborane (volume approximately that swept out by a rotating phenyl ring) we might obtain a close structural analogue of MOF-5, but with smaller pores and therefore higher heats of adsorption. Such a change could lead to higher hydrogen loading at low temperatures and modest pressures and/or more persistent loading at higher temperatures.¹² The solvothermal synthesis instead yielded a more complex MOF of the formula [Zn₃(OH)(*p*-CDC)_{2.5}(DEF)₄]_{*n*} (**2**) [DEF = diethylformamide] (Figure 1). Compound **2** deviates from MOF-5 because of the sterics imposed by the cage and/or the source of the zinc ion used.¹³ X-ray analysis of a single crystal of **2** revealed a structure in which two of the three zinc ions are coordinated to two DEF molecules, each

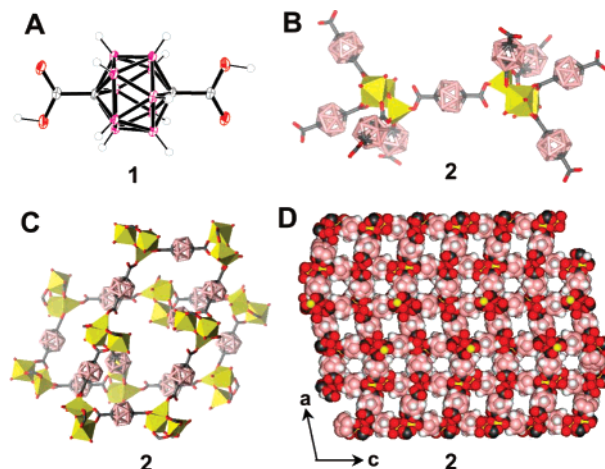


Figure 1. Crystallographically derived (a) structure of **1**, (b) topology and connectivity of **2**, (c) 3-D topology and connectivity of **2**, (d) space-filling packing diagrams of **2** down *b* axes with DEF removed from Zn₃OH clusters. The coordinated DEF molecules are omitted for clarity (b–d).

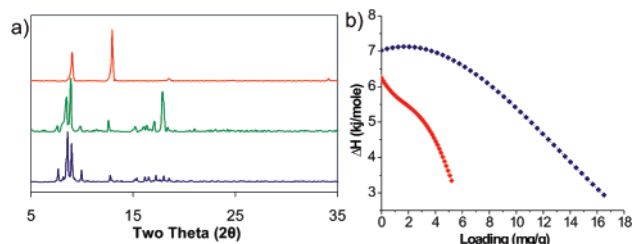


Figure 2. (a) PXRD patterns: simulated pattern based on the single-crystal structure of **2** (bottom), **2** as synthesized (middle), and **4** (top); (b) heats of adsorption (ΔH_{ads}) for H₂ in **4** (blue) and **3** (red).

in an octahedral geometry. In addition, one of the dicarboxylate ligands in the structure is ligated to zinc through only one oxygen atom. The zincs are further connected by a triply bridging hydroxide ion.¹⁴ Thermogravimetric analysis (TGA) of **2** revealed mass losses between 125 and 175 °C and 175–250 °C, assigned to free and coordinated DEF, respectively, but no further mass loss up to 500 °C. Elemental analysis measurements of the crystalline material heated under vacuum at 300 °C confirmed the removal of the coordinated DEF. Powder X-ray diffraction (PXRD) measurements established that although the crystallinity is retained, the structure is irreversibly altered. We have not yet been able to obtain a single-crystal structure of the DEF-free version of the MOF (**4**) (Figure 2a). Infrared data are inconclusive regarding whether the partially coordinated carboxylate of **2** becomes fully coordinated in **4** (Supporting Information, Figure S9). Nonetheless, the number of coordination sites occupied by DEF in **2** is greater than the number

[†] Northwestern University.

[‡] Argonne National Laboratory.

[§] University of Missouri-Columbia.

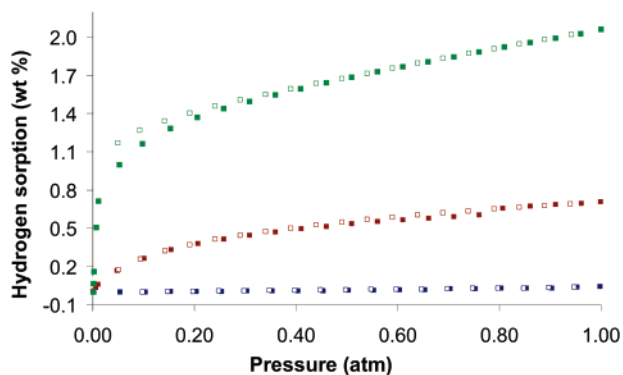


Figure 3. Adsorption (solid squares) and desorption (empty squares) isotherms for H₂ uptake by chloroform-treated **2** (blue), **3** (red), and **4** (green) at 77 K.

of coordination sites needed for complete coordination of *p*-CDC²⁻. This mismatch may result in coordinatively unsaturated or at least highly reactive metal sites.¹⁵ Adsorption measurements (Figure S6) were used to determine the N₂-accessible surface areas of **3** (compound **2** evacuated at 100 °C) and **4**. The Brunauer, Emmett, and Teller (BET) surface areas are 248 and 152 m²/g, respectively. Notably, despite the removal of coordinated DEF, the conversion of **3** to **4** decreases the size of the most prevalent pores from 6 to 5 Å (Figure S10).

In contrast to modest N₂ adsorption, **4** displays remarkably high H₂ uptake at 77 K: 2.1% (25 mg of H₂/cm³ of **4**) at 1 atm. This is apparently triple the uptake of **3** (volumetric measurements; Figure 3) and stands in striking contrast to unheated samples of **2** (i.e., no uptake even after one week of exposure to chloroform (often effective for solvent exchange) and ambient vacuum). Further comparison of **3** and **4** reveals that at 1 atm the latter takes up about 6 additional H₂ molecules per [Zn₃(OH)]⁺⁵ cluster. It is tempting to attribute the enhanced uptake to putative open metal sites, although the observed reduction of pore size might also account for the findings.¹⁶ The H₂ results for **4** compare favorably to those for a variety of other MOFs measured under the same conditions and indeed are exceeded by only three other framework materials,^{7a–c} the highest being 2.45%.^{7a}

Isosteric heats of adsorption for H₂ in **3** and **4** were obtained by fitting 77 and 87 K isotherms to appropriate virial equations (Figure 2b).¹⁷ The Δ*H*_{ads}(H₂) values for **4** are substantially higher than for **3** over the entire loading range. These results are likewise qualitatively consistent with effects expected from reduction of pore size and/or formation of open metal-coordination sites. To summarize, the first carborane-based MOF (**2**) has been synthesized. Removal of the coordinated solvent molecules triples the uptake of H₂ by the material at 77 K and 1 atm (**4** vs **3**) despite a decrease in surface area and reduction of N₂-accessible pore volume. The resistance to pore collapse upon conversion of **3** to **4** likely reflects the rigidity of the dicarborane, its 3-dimensional sterics, and its inability to benefit greatly from stacking-type (collapsed structure) van der Waals interactions. Work on related carborane-based MOFs and on chemical separations using **4** is in progress.

Acknowledgment. J.T.H. gratefully acknowledges the Office of Science, U.S. Dept. of Energy (Grant No. DE-FG02-01ER15244), Argonne Natl. Lab (Fellowship for K.L.M.) and the NU NSF-NSEC for financial support. C.A.M. acknowledges the ARO, AFOSR, and NIH (Director's Pioneer Award). M.F.H. acknowledges support from the NSF.

Supporting Information Available: Experimental procedures for the synthesis of *p*-CDCH₂, **2** and **4**; X-ray crystallographic files for *p*-CDCH₂ and **2** in CIF format; TGA, FTIR, N₂ isotherms and heat of adsorption plots of **2** and **4**. This material is available free of charge via the Internet at <http://pubs.acs.org>.

References

- (1) Oh, M.; Mirkin, C. A. *Nature* **2005**, *438*, 651–654.
- (2) Min, K. S.; Suh, M. P. *J. Am. Chem. Soc.* **2000**, *122*, 6834–6840.
- (3) (a) Cho, S.-H.; Ma, B.-Q.; Nguyen, S. T.; Hupp, J. T.; Albrecht-Schmitt, T. E. *Chem. Commun.* **2006**, 2563–2565. (b) Wu, C. D.; Hu, A.; Zhang, L.; Lin, W. *J. Am. Chem. Soc.* **2005**, *127*, 8940–8941. (c) Gomez-Lor, B.; Gutierrez-Puebla, E.; Iglesias, M.; Monge, M. A.; Ruiz-Valero, C.; Snejko, N. *Chem. Mater.* **2005**, *17*, 2568–2573. (d) Kitaura, R.; Onoyama, G.; Sakamoto, H.; Matsuda, R.; Noro, S. I.; Kitagawa, S. *Angew. Chem., Int. Ed.* **2004**, *43*, 2684–2687.
- (4) Evans, O. R.; Lin, W. *Chem. Mater.* **2001**, *13*, 2705–2712.
- (5) (a) Lee, E. Y.; Jang, S. Y.; Suh, M. P. *J. Am. Chem. Soc.* **2005**, *127*, 6374–6381. (b) Dinca, M.; Long, J. R. *J. Am. Chem. Soc.* **2005**, *127*, 9376–9377.
- (6) Zhao, B.; Chen, X. Y.; Cheng, P.; Liao, D. Z.; Yan, S. P.; Jiang, Z. H. *J. Am. Chem. Soc.* **2004**, *126*, 15394–15395.
- (7) (a) Liu, Y.; Eubank, J. F.; Cairns, A. J.; Eckert, J.; Kravtsov, V. C.; Luebke, R.; Eddaoudi, M. *Angew. Chem., Int. Ed.* **2007**, *46*, 3278–3283. (b) Chen, B.; Ockwig, N. W.; Millard, A. R.; Contreras, D. S.; Yaghi, O. M. *Angew. Chem., Int. Ed.* **2005**, *44*, 4745–4749. (c) Dinca, M.; Dailly, A.; Liu, Y.; Brown, C. M.; Neumann, D. A.; Long, J. R. *J. Am. Chem. Soc.* **2006**, *128*, 16876–16883. (d) Choi, H. J.; Suh, M. P. *J. Am. Chem. Soc.* **2004**, *126*, 15844–15851. (e) Bradshaw, D.; Claridge, J. B.; Cussen, E. J.; Prior, T. J.; Rosseinsky, M. J. *Acc. Chem. Res.* **2005**, *38*, 273–282. (f) Kitagawa, S.; Kitaura, R.; Noro, S. *Angew. Chem., Int. Ed.* **2004**, *43*, 2334–2375. (g) Latroche, M.; Surblé, S.; Serre, C.; Mellot-Draznié, C.; Llewellyn, P. L.; Lee, H.; Chang, J.; Jung, S. H.; Férey, G. *Angew. Chem., Int. Ed.* **2006**, *45*, 8227–8231. (h) Mulfort, K. L.; Hupp, J. T. *J. Am. Chem. Soc.* **2007**, *129*, 9604–9605.
- (8) However, the use of phenyl boronic acids as building blocks for high porosity, metal-free, covalent organic frameworks has recently been reported, see: El-Kaderi, H. M.; Hunt, J. R.; Mendoza-Cortés, J. L.; Côté, A. P.; Tylor, R. E.; O'Keeffe, M.; Yaghi, O. M. *Science* **2007**, *316*, 268–272.
- (9) Hawthorne, M. F. *Advances in Boron Chemistry*; Special Publication No. 201; Royal Society of Chemistry: London, 1997; Vol. 82, p 261.
- (10) (a) Hawthorne, M. F.; Farha, O. K.; Julius, R.; Ma, L.; Jalisatgi, S. S.; Li, T.; Bayer, M. J. *ACS Symp. Ser.* **2006**, *917*, 312–324. (b) Julius, R. L.; Farha, O. K.; Chiang, J.; Perry, L. J.; Hawthorne, M. F. *Proc. Natl. Acad. Sci. U.S.A.* **2007**, *104*, 4808–4813.
- (11) Panella, B.; Hirscher, M. *Adv. Mater.* **2005**, *17*, 538–541.
- (12) Pan, L.; Parker, B.; Huang, X.; Olson, D. H.; Lee, J. Y.; Li, J. *J. Am. Chem. Soc.* **2006**, *128*, 4180–4181.
- (13) (a) Hausdorf, S.; Baitalow, F.; Seidel, J.; Mertens, F. *J. Phys. Chem. A* **2007**, *111*, 4259–4266. (b) Liao, J. H.; Lee, T. J.; Su, C. T. *Inorg. Chem. Commun.* **2006**, *9*, 201–204. (c) Loiseau, T.; Muguerra, H.; Férey, G.; Haouas, M.; Taulelle, F. *J. Solid State Chem.* **2005**, *178*, 621–628.
- (14) For another example of such bonding, see: Vodak, D. T.; Braun, M. E.; Kim, J.; Eddaoudi, M.; Yaghi, O. M. *Chem. Commun.* **2001**, 2534–2535.
- (15) See for example: Moon, H. R.; Kobayashi, K.; Suh, M. P. *Inorg. Chem.* **2006**, *45*, 8672–8676.
- (16) Jung, S. H.; Kim, H. K.; Yoon, J. W.; Chang, J. S. *J. Phys. Chem. B* **2006**, *110*, 9371–9374.
- (17) Czepirski, L.; Jagiello, J. *Chem. Eng. Sci.* **1989**, *44*, 797–801.

JA076167A